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provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2.

6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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is the key to nearly
everything It falls from
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ANSWER Answer the
following questions in
the space provided. 1.

Given the following
equation: $C_3H_4(g) + x$
 $O_2(g) \rightarrow 3CO_2(g) +$

$2H_2O(g)$ a. What is the
value of the coefficient
. x. in this equation? b.

What is the molar mass

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Answers

of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ? 2. a. What is meant by . ideal conditions

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1-3 Introduction to

Stoichiometry Ideal

Stoichiometric

Calculations Limiting

Reactant and Percent

Yield 2 Definitions.

Composition

stoichiometry deals

with the mass

relationships of

elements in

compounds. Reaction

stoichiometry involves

the mass

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Chemistry Chapter 9
Stoichiometry
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SECTION 9.2.

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$

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Stoichiometry
Modern Chemistry
Answers

$2\text{KClO}_3(\text{s}) + 3\text{O}_2(\text{g})$ How many grams of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2. Given the following equation ...

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made through "File
info" ... Reaction
stoichiometry, the
subject of this chapter,

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is based on chemical equations and the law of conservation of mass. All reaction stoichiometry calculations start with a balanced chemical equation. This equation gives the

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Reactions of Alkenes
Review of Concepts Fill
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verify that your
answers are correct,
look in your textbook
at the end of Chapter
9. Each of the
sentences below
appears verbatim in
the section entitled
Review of Concepts
and Vocabulary.

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the left Show all your

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work in the space provided 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 30 mol ...

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one can use
stoichiometry to
calculate masses,
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STOICHIOMETRY The following pages contain the bulk (but not all) of the information for the chapter 9 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.

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